##  **Stage 2 Chemistry**

##  **Birdwood**

 HIGH SCHOOL **Topic 3: Using and Controlling Reactions**

 **Chemical Industry and Metal Production**

 **Review Paper 17**

**DUE DATE:** Ref: ESSENTIALS pages 179 - 196

**Question 1**

The conversion of nitrogen and hydrogen to ammonia in the Haber Process is represented by the following equation:

 N2(g) + 3H2(g) 2NH3(g) ΔH = -92 kJmol-1.

A catalyst and a relatively low temperature are used in this reaction to give the best yield.

a) State the *purpose* of the catalyst in this reaction.

b) The beginning of the energy profile

diagram for the reaction is shown opposite.

 Complete the energy profile diagram,

 clearly identifying the ΔH and the

 activation energy.

c) On the energy profile diagram, draw the

 reaction pathway of this reaction

 without the catalyst.

d) A compromise between yield and rate of reaction often determines the optimum conditions for an industrial process.

 i State and explain the effect of low temperature on the yield of ammonia.

 ii In terms of Collision Theory, state and explain the effect low temperature has on the rate of this

 reaction.

The table below gives the percentages of ammonia in the equilibrium mixture at various pressures.

|  |  |
| --- | --- |
| **Pressure (atmospheres)** | **Yield of ammonia at 5500C with a catalyst** |
| 11001000 | negligible7%41% |

In industry, a pressure of 250 atmospheres is used in the above equilibrium.

e) State one reason why a pressure of 250 atmospheres is used and not 1000 atmospheres.

In one experiment conducted at 5500C, the concentrations of all three gases in the system were determined. The results are shown in the table below:

|  |  |
| --- | --- |
| **Gas** | **Concentration****(mol L-1)** |
| N2H2NH3 | 0.50.30.3 |

f) Write a Kc expression for this reaction.

g) The value for Kc for this reaction at 5500C is **0.35**.

 Using your Kc expression, *calculate the value of Kc*, using the values in the above table, to explain why the

 system is not at equilibrium.

 (16 marks)

**Question 2**

In the production of zinc, the zinc sulfide ore is roasted in air to form zinc oxide.

Sulfur dioxide gas is the other product, which is used in the manufacture of sulfuric acid. The sulfuric acid is then reacted with zinc oxide in a leaching process to produce a solution of zinc sulfate.

The zinc sulfate solution must be purified before electrolysis.

a) Write an equation to represent the roasting stage.

*The next stage is the leaching of zinc oxide with sulfuric acid.*

b) Write an equation to represent this stage.

*Zinc sulfate is the electrolysed to produce zinc metal.*

c) Write a half-equation to represent the formation of zinc.

d) Does this half-equation represent oxidation or reduction?

e) Will this occur at the positive or negative electrode of the electrolysis unit?

*The purification process is essential in the production of zinc.*

f) Describe the process that purifies the Zn2+(aq).

g) State how the final product (zinc) after electrolysis, would be different if purification is not done.

*Environmentalists and conservationists will often view mining and metal production in a critical way.*

h) Describe *one harmful effect* zinc mining or zinc metal production may have on the environment and

 *one method to reduce this effect.*

 (12 marks)

**Question 3**

Magnesium is produced by a series of processes as shown in the flowchart below:

 CaCO3(S)

heater

 CaO(s)

slaker

 H2O(l)

 Ca(OH)2(aq)

precipitator

 sea water

 Mg(OH)2(s)

neutraliser

 HCl

 MgCl2(aq)

acid

plant

solar evaporators

driers

 MgCl2(s)

electrolysis unit

 Cl2

 **Magnesium**

a) Name the *three* raw materials used in this process.

b) Name the *by-product* obtained from the electrolysis unit.

c) What use is made of the by-product and how might this increase the profitability of the company?

d) From the flowchart, identify *two* stages at which large amounts of energy are needed.

e) Complete the chemical equation for the reaction occurring in the precipitator:

 Ca(OH)2(aq) + Mg2+(aq)

f) Name the other substance produced when calcium carbonate is converted into calcium oxide in the heater.

g) Why might this part of the process be of concern to environmentalists?

In the electrolysis unit, *molten* magnesium chloride is electrolysed and magnesium is produced at one of the electrodes.

h) Explain why all of the water must be removed from the magnesium chloride before it is electrolysed.

i) Write a half equation for the electrode reaction in which magnesium is produced.

j) Name the electrode at which the magnesium is produced.

 (16 marks)

**Question 4**

a) What changes will always shift this equilibrium reaction to the right?

2HI(*g*) → H2(*g*) + I2(*g*) Δ*H* = –52 kJ

b) Consider the following mixture of gases in a closed 5.0 L vessel at 730°C.

 The following reaction occurs:

 CH4(*g*) +H2O(*g*) → CO(*g*) +3H2(*g*) ∆*H*  = -206 kJ

 The equilibrium constant, K, is 0.26 at 730°C.

1. Determine whether the system is at equilibrium.
2. Explain how conditions in this reaction could be adjusted to increase the quantity of products.

(9 marks)

 **TOTAL MARK = 53**